

Sr. No.	Client Question ID	Question Body and Alternatives	Marks	Negative Marks
Objective Question				
1	1	What is the maximum frequency of X-ray produced by 30 kV electron? A1 : 7.24 x 10¹⁸ Hz A2 : 7.24 x 10²⁶ Hz A3 : 4.80 x 10¹⁸ Hz A4 : 4.80 x 10²⁶ Hz	4.0	1.00
Objective Question				
2	2	The total energy operator is given by A1 : $\hat{E} = i\hbar \frac{\partial}{\partial t}$ A2 : $\hat{E} = i\hbar \frac{d}{dt}$ A3 : $\hat{E} = i\hbar \frac{\partial}{\partial x}$ A4 : $\hat{E} = i\hbar \frac{d}{dx}$	4.0	1.00
Objective Question				
3	3	Splitting of spectrum lines in a magnetic fields is known as A1 : Stark effect A2 : Zeeman effect A3 : Voigt effect A4 : Faraday effect	4.0	1.00
Objective Question				
4	4	The simplified form of $(A+B)(\bar{A}+B)\bar{B}$ Boolean expression is	4.0	1.00

		A1 : 0 A2 : A+B A3 : (A+B) $\bar{B}$ A4 : (A+B) $\bar{A}$		
Objective Question				
5	5	Debye theory of specific heat is valid at A1 : Room Temperature A2 : low temperature A3 : intermediate temperature A4 : all temperature	4.0	1.00
Objective Question				
6	6	Which plane will be present in diffraction pattern of a body centred cubic A1 : (1 0 0) A2 : (1 1 0) A3 : (1 1 1) A4 : (0 1 0)	4.0	1.00
Objective Question				
7	7	The Fermi-Dirac distribution function n(E)] is (k = Boltzmann constant, T=temperature and E_f = Fermi energy) A1 : $n(E) = \frac{1}{e^{\frac{E-E_f}{kT}} + 1}$ A2 : $n(E) = \frac{1}{e^{\frac{E_f-E}{kT}} + 1}$ A3 : $n(E) = \frac{1}{e^{\frac{E-E_f}{kT}} - 1}$ A4	4.0	1.00

		: $n(E) = \frac{1}{e^{kT} \frac{E+E_f}{-1}}$		
Objective Question				
8	8	A sample germanium with intrinsic carrier density if $9.65 \times 10^9 \text{ cm}^{-3}$ is doped with 10^{16} phosphorous atoms/cm^3, then Hall co-efficient is A1 : $625 \text{ cm}^3 \text{ C}$ A2 : $-965 \text{ cm}^3 / \text{C}$ A3 : $-625 \text{ cm}^3 / \text{C}$ A4 : $965 \text{ cm}^3 / \text{C}$	4.0	1.00
Objective Question				
9	9	How many fermions can be adjusted into three dimensional harmonic oscillator upto energy levels $\frac{7}{2} \hbar \omega$ A1 : 8 electron A2 : 12 electron A3 : 16 electron A4 : 20 electron	4.0	1.00
Objective Question				
10	10	The entropy S is proposal to (Ω is number of microstates) A1 : $S = \frac{k_B}{\ln \Omega}$ A2 : $S = \frac{\ln \Omega}{k_B}$ A3 : $S = k_B \ln \Omega$ A4 : $S = k_B \Omega$	4.0	1.00
Objective Question				
11	11	The nuclear spin and parity of ${}^{40}_{20}\text{Ca}$ in its ground state is A1 : 0^+	4.0	1.00

		A2 1^+ : A3 0^- : A4 1^- :		
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Objective Question

12	12	For the adiabatic expansion of a blackbody radiation enclosure, which of the following is correct? A1 $V.T = \text{constant}$: A2 $V/T = \text{constant}$: A3 $V^{1/3} T = \text{constant}$: A4 $V^{4/3} T = \text{constant}$:	4.0	1.00
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Objective Question

13	13	The Eigen values of an observable associated with a physical quality are 0 and 1. The expectation values of measurements is A1 0 : A2 Either 0 or 1 : A3 between 0 and 1 : A4 1 :	4.0	1.00
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Objective Question

14	14	A phosphorous doped silicon semiconductor (doping density: $10^{17}/\text{cm}^3$) is heated from 100 °C to 200 °C. Which one of the following statements is CORRECT? A1 Position of Fermi level moves towards conduction band : A2 Position of dopant level moves towards conduction band : A3 Position of Fermi level moves towards middle of energy gap : A4 Position of dopant level moves towards middle of energy gap :	4.0	1.00
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Objective Question

15	15	The magnetic susceptibility is negative for	4.0	1.00
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		A1 : paramagnetic material only A2 : ferromagnetic material only A3 : both paramagnetic and ferromagnetic materials A4 : diamagnetic material only		
Objective Question				
16	16	The zeroth law of thermodynamics allows us to define: A1 : temperature A2 : pressure A3 : work A4 : internal energy	4.0	1.00
Objective Question				
17	17	The de Broglie wavelength of an electron of energy 200 MeV is A1 : $0.62 \times 10^{-8} \text{ m}$ A2 : $0.62 \times 10^{-10} \text{ m}$ A3 : $0.62 \times 10^{-12} \text{ m}$ A4 : $0.62 \times 10^{-14} \text{ m}$	4.0	1.00
Objective Question				
18	18	Unit of reduced Planck's constant ($\hbar$) is A1 : J.s^{-1} A2 : $\text{Kg.m}^2. \text{s}^{-1}$ A3 : J.s^1 A4 : $\text{Kg.m}^2. \text{s}^{-2}$	4.0	1.00
Objective Question				

19	19	The Ampere's law with Maxwell's correction term is A1 : $\nabla \times \vec{B} = \mu_0 \vec{J} + \mu_0 \epsilon_0 \frac{\partial \vec{E}}{\partial t}$ A2 : $\nabla \times \vec{B} = \mu_0 \vec{J} + \mu_0 \epsilon_0 \frac{\partial \vec{B}}{\partial t}$ A3 : $\nabla \times \vec{E} = \mu_0 \vec{J} + \mu_0 \epsilon_0 \frac{\partial \vec{E}}{\partial t}$ A4 : $\nabla \times \vec{E} = \mu_0 \vec{J} + \mu_0 \epsilon_0 \frac{\partial \vec{B}}{\partial t}$	4.0	1.00
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Objective Question

20	20	Dimensional formula of solid angle is A1 : $[M L^2 T^0]$ A2 : $[M L^2 T^{-1}]$ A3 : $[M L^0 T^{-2}]$ A4 : Dimensionless	4.0	1.00
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Objective Question

21	21	How many gram of H_2SO_4 are present in 0.25 mole of H_2SO_4 A1 : 2.45 A2 : 24.5 A3 : 0.245 A4 : 0.250	4.0	1.00
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Objective Question

22	22	Which of the following does not contain co-ordinate bond A1 : BH_4^- A2 : NH_4^+ A3 : CO_3^{2-} A4 : H_3O^+	4.0	1.00
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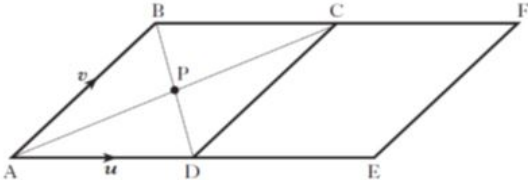
Objective Question				
23	23	According to Fajan's rule covalent bond is favoured by A1 : Large cation and small anion A2 : Large cation and large anion A3 : Small cation and small anion A4 : Small cation and large anion	4.0	1.00
Objective Question				
24	24	Heat of combustion is always A1 : +ve A2 : -ve A3 : Neutral A4 : All of these	4.0	1.00
Objective Question				
25	25	The osmotic pressure of a dilute solution is directly proportional to the A1 : Diffusion rate of solute A2 : Ionic concentration A3 : Boiling point A4 : Flow of solvent from a concentrated to a dilute solution	4.0	1.00
Objective Question				
26	26	The vapour density of completely dissociated NH_4Cl would be A1 : Same as that of NH_4Cl A2 : Double than that of NH_4Cl A3 : Half than that of NH_4Cl	4.0	1.00

		A4 Slightly less than that of NH_4Cl :		
Objective Question				
27	27	What is the conjugate base of OH^- ? A1 O_2 : A2 H_2O : A3 O^- : A4 O^{2-} :	4.0	1.00
Objective Question				
28	28	In which of the following states the particle size would be greater than 1000 nm? A1 Suspension : A2 True solution : A3 Colloidal Solution : A4 None of these :	4.0	1.00
Objective Question				
29	29	Collosion theory is applicable to A1 First order reaction : A2 Zero order reaction : A3 Bimolecular reaction : A4 Intramolecular reaction :	4.0	1.00
Objective Question				
30	30	The Oxidation number of Fe in $\text{K}_3[\text{Fe}(\text{CN})_6]$ is A1 +2 : A2 +3 :	4.0	1.00

		A3 +1 : A4 +4 :		
Objective Question				
31	31	The unit of equivalent conductivity is A1 Ohm cm : A2 $\text{Ohm}^{-1}\text{cm}^2(\text{g equivalent})^{-1}$: A3 $\text{Ohm cm}^2 (\text{g equivalent})$: A4 Scm^{-2} :	4.0	1.00
Objective Question				
32	32	H_2O_2 has reducing action upon A1 PbS : A2 O_3 : A3 H_2S : A4 Na_2SO_3 :	4.0	1.00
Objective Question				
33	33	The substance not likely to contain CaCO_3 is A1 Calcined gypsum : A2 Sea shells : A3 Dolomite : A4 A marble statue :	4.0	1.00
Objective Question				
34	34	Which is in correct about H_2SO_4 ? A1 Reducing agent :	4.0	1.00

		A2 Dehydrating agent : A3 Sulphonating agent : A4 Highly viscous :		
Objective Question				
35	35	Which is paramagnetic? A1 $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$: A2 $[\text{Fe}(\text{CN})_6]^{4-}$: A3 $[\text{Ni}(\text{CO})_4]$: A4 $[\text{Ni}(\text{CN})_4]^{2-}$:	4.0	1.00
Objective Question				
36	36	Turpentine oil can be purified by A1 Vacuum distillation : A2 Fractional distillation : A3 Steam distillation : A4 Simple distillation :	4.0	1.00
Objective Question				
37	37	Which of the following is strongest base in aqueous solution? A1 Methylamine : A2 Trimethylamine : A3 Aniline : A4 Dimethylamine :	4.0	1.00
Objective Question				
38	38	A ketoxime on reduction and followed by acetylation gives _____ A1 Ethylamine	4.0	1.00

		: A2 Isopropyl amine : A3 Monoacetyl isopropyl amine : A4 Diacetyl isopropyl amine :		
Objective Question				
39	39	Which of the following is not a nucleophile? A1 CN^- : A2 OH^- : A3 NH_3 : A4 BF_3 :	4.0	1.00
Objective Question				
40	40	Chloroform reaction with concentrated HNO_3 gives A1 Chloropicrin : A2 Picric acid : A3 Acetylene : A4 Nitromethane :	4.0	1.00
Objective Question				
41	41	$(2^{\log_{1/3} \frac{1}{8}})^{\log_4 3}$ equals A1 1 : A2 2 : A3 $2\sqrt{2}$: A4 4 :	4.0	1.00

Objective Question			
42	42	Each vertex of a cube is assigned an integer so that no two vertices connected by an edge of the cube are assigned the same integer. The least number of integers required for such an assignment is A1 2 : A2 3 : A3 4 : A4 6 :	4.0 1.00
Objective Question			
43	43	The diagram shows two congruent parallelograms. Vectors u and v are represented by AD and AB.  Express EP in terms of u and v. A1 $-\frac{3}{2}u + \frac{1}{2}v$: A2 $\frac{3}{2}u - \frac{1}{2}v$: A3 $-\frac{3}{2}u - \frac{1}{2}v$: A4 $\frac{3}{2}u + \frac{1}{2}v$:	4.0 1.00
Objective Question			
44	44	P, Q and R have coordinates $(1, -2, 4)$, $(5, 4, -2)$ and $(7, 7, -5)$ respectively. Here are two statements about the points P, Q and R: (1) P, Q and R are collinear (2) $PR : QR = 2:1$ Which of the following is true? A1 Neither of the statement are correct : A2 Only statement (1) is correct : A3 Only statement (2) is correct :	4.0 1.00

		A4 Both statements are correct :		
Objective Question				
45	45	Functions f and g are given by $f(x) = 3x^2 - 1$ and $g(x) = x^2 + 2$. Find an expression for $f(g(x))$ A1 $4x^2 + 1$: A2 $9x^4 + 1$: A3 $3x^4 + 5x^2 - 2$: A4 $3x^4 + 12x^2 + 11$:	4.0	1.00
Objective Question				
46	46	A circle with centre $(-1, 5)$ passes through the point $(2, 1)$. What is the equation of the circle? A1 $(x + 1)^2 + (y - 5)^2 = 37$: A2 $(x - 1)^2 + (y - 5)^2 = 37$: A3 $(x + 1)^2 + (y - 5)^2 = 25$: A4 $(x - 1)^2 + (y + 5)^2 = 25$:	4.0	1.00
Objective Question				
47	47	If the system of linear equations $x + 2ay + az = 0$ $x + 3by + bz = 0$ $x + 4cy + cz = 0$ has a non-zero solution, then a, b, c A1 are in A. P. : A2 are in H. P. : A3 are in G.P. : A4 satisfy $a + 2b + 3c = 0$:	4.0	1.00
Objective Question				
48	48	The area of the region bounded by the curves $y = x - 1 $ and $y = 3 - x $ is	4.0	1.00

		A1 2 sq units : A2 3 sq units : A3 4 sq units : A4 6 sq units :		
Objective Question				
49	49	Let A (2, -3) and B(-2, 1) be vertices of a triangle ABC. If the centroid of this triangle moves on the line $2x + 3y = 1$, then the locus of the vertex C is the line A1 $2x - 3y = 7$: A2 $3x + 2y = 5$: A3 $3x - 2y = 3$: A4 $2x + 3y = 9$:	4.0	1.00
Objective Question				
50	50	Suppose f(x) is differentiable at $x = 1$ and $\lim_{h \rightarrow 0} \frac{1}{h} f(1+h) = 5$, then $f'(x) =$ A1 3 : A2 4 : A3 5 : A4 6 :	4.0	1.00
Objective Question				
51	51	If $2a + 3b + 6c = 0$, then at least one root of the equation $ax^2 + bx + c = 0$ lies in the interval A1 (0, 1) : A2 (1, 2) : A3 (2, 3) : A4 (1, 3)	4.0	1.00

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Objective Question				
52	52	If y is a twice differentiable function and $x \cos y + y \cos x = \pi$, then $y''(0) =$ A1 0 : A2 1 : A3 $-\pi$: A4 π :	4.0	1.00
Objective Question				
53	53	If $f(x) = x(x - x+1)$, then A1 $f(x)$ is continuous but not differentiable at $x = 0$: A2 $f(x)$ is differentiable at $x = 0$: A3 $f(x)$ is not differentiable at $x = 0$: A4 none of these :	4.0	1.00
Objective Question				
54	54	The value of $\lim_{x \rightarrow 0} \frac{\sqrt{\frac{1}{2}(1 - \cos 2x)}}{x}$ is A1 1 : A2 -1 : A3 0 : A4 none of these :	4.0	1.00
Objective Question				
55	55	The area enclosed between the curve $y = \log_e (x+e)$ and the coordinate axes is A1 1 : A2 2 :	4.0	1.00

		A3 3 : A4 4 :		
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Objective Question

56	56	The value of $\int_{-2}^3 1-x^2 dx$ is A1 28/3 : A2 1/3 : A3 7/3 : A4 4 :	4.0	1.00
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Objective Question

57	57	$\int_0^1 \sqrt{\frac{1-x}{1+x}} dx$ is equal to A1 $\frac{\pi}{2} + 1$: A2 $\frac{\pi}{2} - 1$: A3 1 : A4 π :	4.0	1.00
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Objective Question

58	58	The value of $\int \frac{\log x}{x^2} dx$ is A1 $\log(x+1) + c$: A2 $-\frac{1}{x} \log(x+1) + c$: A3 $\log(x-1) + c$: A4 $\frac{1}{2} \log(x+1) + c$:	4.0	1.00
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Objective Question

59	59		4.0	1.00
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		If $f(x) = \int_1^x \sqrt{2-t^2} dt$, then the real roots of the equation $x^2 - f'(x) = 0$ are A1 ± 1 : A2 $\pm \frac{1}{\sqrt{2}}$: A3 $\pm \frac{1}{2}$: A4 0 and 1 :		
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Objective Question				
60	60	The slope of the tangent to a curve $y = f(x)$ at $[x, f(x)]$ is $2x+1$. If the curve passes through the point (1,2), then the area of the region bounded by the curve, the X-axis and the line $x = 1$ is A1 6 : A2 $6/5$: A3 1 : A4 $5/6$:	4.0	1.00

Objective Question				
61	61	If a cell is immersed in a concentrated solution, it follow the phenomenon A1 Turgor : A2 Plasmolysis : A3 Hemolysis : A4 Paralysis :	4.0	1.00

Objective Question				
62	62	Plasma membrane is made up of A1 Protein, lipid, carbohydrate : A2 Lipid, carbohydrate : A3 Protein, lipid :	4.0	1.00

		A4 Protein :		
Objective Question				
63	63	Which one of the following organisms is not an example of eukaryotic cells	4.0	1.00
		A1 Amoeba proteus :		
		A2 Paramecium caudatum :		
		A3 Escherichia coli :		
		A4 Euglena viridis :		
Objective Question				
64	64	Which one of the following pairs of nitrogenous bases of nucleic acids, is wrongly matched with the category mentioned against it?	4.0	1.00
		A1 Adenine, Thymine – Purines :		
		A2 Thymine, Uracil – Pyrimidines :		
		A3 Uracil, Cytosine – Pyrimidines :		
		A4 Guanine, Adenine – Purines :		
Objective Question				
65	65	The two polynucleotide chains in DNA are:	4.0	1.00
		A1 Semiconservative :		
		A2 Parallel :		
		A3 Discontinuous :		
		A4 Antiparallel :		
Objective Question				
66	66	Triacylglycerols are	4.0	1.00
		A1 soluble in water :		
		A2 insoluble in water :		

		A3 soluble in water at elevated temperature : A4 partially soluble in water :		
Objective Question				
67	67	Hydrolysis of lactose yields A1 galactose and fructose : A2 galactose and glucose : A3 glucose and fructose : A4 fructose and galactose :	4.0	1.00
Objective Question				
68	68	Phospholipid contains A1 hydrophilic heads and hydrophobic tails : A2 long water-soluble carbon chains : A3 positively charged functional groups : A4 both long water-soluble carbon chains and positively charged functional groups :	4.0	1.00
Objective Question				
69	69	In DNA double helix, the two DNA chains are held together by A1 covalent bonds between the pair of bases : A2 hydrogen bonds between the pair of bases : A3 ionic bonds between the pair of bases : A4 none of these :	4.0	1.00
Objective Question				
70	70	Which of the following is incorrect? A1 In DNA double helix, two strands of the DNA are bound with each other with the bases : A2 Adenine always pairs with thymine	4.0	1.00

		: A3 Guanine always pairs with the cytosine : A4 None of these :		
Objective Question				
71	71	Pyruvic acid before entering the Kreb's cycle is changed into A1 Citric acid : A2 Succinic acid : A3 Acetyl CoA : A4 Malic acid :	4.0	1.00
Objective Question				
72	72	A noncompetitive inhibitor of an enzyme-catalyzed reaction A1 increases K_m and increases V_{max} : A2 increases K_m and reduces V_{max} : A3 reduces K_m and increases V_{max} : A4 reduces K_m and reduces V_{max} :	4.0	1.00
Objective Question				
73	73	Which of the following enzyme is used to covalently bond foreign DNA to a vector plasmid? A1 DNA polymerase : A2 Restriction endonuclease : A3 DNA ligase : A4 DNA helicase :	4.0	1.00
Objective Question				
74	74	Which of the following statements is true about transmembrane electrical potential? A1 The chemical gradient drives Cl^- and K^+ inwards :	4.0	1.00

		A2 : The electrical gradient drives Na^+ and Ca^{+2} inwards A3 : The chemical gradient drives Na^+ and Ca^{+2} inwards and K^+ outward A4 : The chemical gradient drives Cl^- and K^+ outwards		
Objective Question				
75	75	Frame shift mutation may occur as a result of A1 : formation of a thymine-dimer A2 : deamination of cytosine to uracil A3 : conversion of guanine to xanthine A4 : insertion/deletion of a nucleotide in which the number of deleted base pairs is not divisible by three	4.0	1.00
Objective Question				
76	76	COVID-19 is A1 : DNA-Virus A2 : RNA-Virus A3 : Bacteria A4 : Blue Green Algae	4.0	1.00
Objective Question				
77	77	A totipotent cell means A1 : An undifferentiated cell capable of developing into a system or entire plant A2 : An undifferentiated cell capable of developing into an organ A3 : An undifferentiated cell capable of developing into complete embryo A4 : Cell which lacks the capability differentiate into an organ or system	4.0	1.00
Objective Question				
78	78	Reverse transcriptase produces	4.0	1.00

		A1 DNA from peptides : A2 RNA from DNA : A3 RNA from RNA : A4 DNA from RNA :		
Objective Question				
79	79	Zinc finger are characteristics of A1 Blood clotting proteins : A2 DNA claperones : A3 DNA binding proteins : A4 Lysosomal hydrolases :	4.0	1.00
Objective Question				
80	80	Average diameter of COVID-19 particle in A1 micro meter : A2 nano meter : A3 feet : A4 inches :	4.0	1.00
Objective Question				
81	81	The Charpy V notch test provides information on A1 the energy absorbed by a material during fracture : A2 the energy absorbed during creep : A3 hardness : A4 the fatigue failure :	4.0	1.00
Objective Question				

82	82	Hume-Rothery rules are associated with A1 : phase diagram A2 : superconductors A3 : thermal expansion A4 : solid solubility	4.0	1.00
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Objective Question

83	83	The dislocation observed in materials is a _____ defect A1 : casting defect A2 : line A3 : plane A4 : point	4.0	1.00
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Objective Question

84	84	Covalent and ionic solids are not suitable as structural components because of their A1 : ductility A2 : brittleness A3 : high elastic moduli A4 : low melting point	4.0	1.00
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Objective Question

85	85	In heat treatment techniques, quenching process refers to A1 : rapid heating A2 : rapid cooling A3 : slow heating A4 : slow cooling	4.0	1.00
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Objective Question				
86	86	The entropy value of pure crystalline material at absolute zero is A1 : 0.01 J/mol/K A2 : 0.1 J/mol/K A3 : 1 J/mol/K A4 : 0 J/mol/K	4.0	1.00
Objective Question				
87	87	Thermal barrier coatings are applied to _____ for protecting them from thermally induced mechanical degradation A1 : turbine components A2 : solid oxide fuel cells A3 : polymers A4 : tensile strength	4.0	1.00
Objective Question				
88	88	Iron alloy containing less than ~2% of carbon is known as A1 : ferrite A2 : cast iron A3 : steel A4 : wrought iron	4.0	1.00
Objective Question				
89	89	The number of degrees of freedom for a solution containing salt dissolved in water A1 : 3 A2 : 2 A3 : 1	4.0	1.00

		A4 0 :		
Objective Question				
90	90	The failure that occurs as a result of alternating cycles of loading and unloading is known as A1 : strain hardening A2 : fatigue A3 : ductile to Brittle transition A4 : creep	4.0	1.00
Objective Question				
91	91	Czochralski method is related to A1 : thin film growth A2 : layer-by-layer growth A3 : island growth A4 : single crystal growth	4.0	1.00
Objective Question				
92	92	The oxide film should be _____ in order to provide a good oxidation resistance. A1 : low in electrical conductivity A2 : high electrical conductivity A3 : continuous A4 : porous	4.0	1.00
Objective Question				
93	93	The highest electrical conductivity among the following is shown by A1 : copper A2 : aluminum A3 silver	4.0	1.00

		: A4 gold :		
Objective Question				
94	94	The process of brazing is associated with A1 forming technique : A2 joining technique : A3 extrusion technique : A4 forging technique :	4.0	1.00
Objective Question				
95	95	Jominy end quench test is associated with the measurement of A1 crack resistance : A2 tensile stress : A3 hardenability : A4 fracture toughness :	4.0	1.00
Objective Question				
96	96	The element which occurs in the metallic form in nature A1 Ag : A2 Au : A3 Al : A4 Fe :	4.0	1.00
Objective Question				
97	97	The ability of a material to absorb energy in the plastic regime is known as A1 elasticity : A2 fatigue :	4.0	1.00

		A3 creep : A4 resilience :		
Objective Question				
98	98	The solubility of hydrogen in steel leads to A1 better weldability : A2 yield strength decrease : A3 embrittlement : A4 photocatalysis :	4.0	1.00
Objective Question				
99	99	The rusting of iron component leads to A1 an increase in weight : A2 a reduction in weight : A3 a constant weight : A4 no change in colour and weight :	4.0	1.00
Objective Question				
100	100	In the vulcanization process, the natural rubber is heated with _____ A1 carbon : A2 silicon : A3 sulphur : A4 phosphorous :	4.0	1.00